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REVIEW SHEET. 1.

What are the 2 main types of bonds that we have learned about? ...

What levels of differences in electronegativity correspond to polar

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covalent, nonpolar  
covalent, and ionic  
bonds? 6. If Sodium,  
Na, was to bond with  
Chlorine, Cl, what type  
of bond would they  
form? ...

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meguarino24. Terms in this set (62) Chemical bond. Mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together. Potential energy.

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Terms in this set (40)  
Bond energy. Energy  
required to break a  
chemical bond. Bond  
length. The distance

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between two bonded atoms at their minimum potential energy, that is, the average between two bonded atoms.

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CHAPTER 6 REVIEW

Chemical Bonding

SECTION 5 SHORT

ANSWER Answer the following questions in the space provided. 1.

Identify the major

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assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear molecule?

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Modern Chemistry 46  
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In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

\_\_\_\_\_ 1. The charge on an ion is a. always positive. b. always negative. c. either positive or negative.

### **Assessment Chapter Test A**

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Practice Test: Chemical Bonding In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

\_\_\_\_\_ 1. The charge on an ion is a. always positive. b. always negative. c.

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SECTION 1 SHORT  
ANSWER Answer the  
following questions in  
the space provided. 1.

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a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

## **Chapter 6 Chemical Bonding Test Answer Key**

Chapter 6: Chemical

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### Chemical Bonding Test Review 1.

List all the symbols of the elements on the periodic table that have a stable electron configuration (aka- a full outer shell). 2. In an electron dot diagram, what does the element symbol represent? 3. Explain when and how ionic bonds form (be specific). 4. Explain when and how covalent bonds form (be specific). 5.

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Bonding DRAFT. ... Two  
or more atoms held  
together by a chemical  
bond is called what?

answer choices . A  
molecule. A molecule.

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15: Total Marks: 30:

Total Time: 20 Minutes:

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Test  
improve their general knowledge or are currently preparing for any ...

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11. Bond energy is the energy

- a. required to break a chemical bond.
- b. released when a chemical bond breaks.
- c. required to form a chemical bond.
- d. absorbed when a chemical bond forms.

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12. In a molecule of fluorine, the two shared electrons give each fluorine atom \_\_\_\_\_ electron(s) in the outer energy level. a. 1 c. 8  
b. 2 d. 32 13. electrons in

**Chapter 12 Review 2**

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