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2.2 MASS-ACTION KINETICS 2.2.1 OVERVIEW We assume that rates can be described by $\text{Rate} = k \cdot (\text{concentration})^n$ where k is the rate constant from Arrhenius' law and n is the rate law. Empirically, we say that $k = A \cdot e^{-E_a/RT}$... which is typical rate law kinetics.

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