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pH/pOH. pH and pOH are a measure of how acidic or how basic a solution is and for most solutions, is measured on a scale of 0 to 14. The lower the number, the more acidic the solution is, and the higher the number, the more basic the solution is. Neutral solutions (not acidic or basic) have a pH of 7.

## pH / pOH - Chemistry Help

pH and pOH CONTINUED 1. 0.01 M HCl  $[H^+] = 0.01M$  pH =  $-\log(0.01) = 2$  2. 0.0010 M NaOH  $[OH^-] = 0.0010M$  pOH =  $-\log(0.010) = 3 \rightarrow$  pH = 11 3. 0.050 M NaOH

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## **pH and pOH CONTINUED - Newbury Park High School**

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pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral 7-14 basic Since  $[H^+][OH^-] = 10^{-14}$  at 25°C, if  $[H^+]$  is known, the  $[OH^-]$  can be calculated and vice versa.  $pH = -\log [H^+]$  So if  $[H^+] = 10^{-6} M$ ,  $pH = 6$   $pOH = -\log [OH^-]$  So if  $[OH^-] = 10^{-8}$ ,  $pOH = 8$  Together,  $pH + pOH = 14$ .

## **pH and pOH - Newbury Park High School**

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Definitions of pH, pOH, and the pH scale. Calculating the pH of a strong acid or base solution. The relationship between acid strength and the pH of a solution.

## **pH, pOH, and the pH scale (article) | Khan Academy**

Acids, Bases, pH and pOH . There are several ways to define acids and bases, but pH and pOH refer to hydrogen ion concentration and hydroxide ion concentration, respectively. The "p" in pH and pOH stands for "negative logarithm of" and is used to make it easier to work with extremely large or small values.

## **Chemistry Review of pOH Calculations**

KEY Chemistry: pH and pOH calculations Part 1: Fill in the missing information in the table below.  $pH [H_3O^+]$   $pOH [OH^-]$  ACID or BASE?  $3.78$   $1.66 \times 10^{-4} M$   $10.22$   $6.03 \times 10^{-11} M$

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Acid 3.41 3.89 x 10<sup>-4</sup> M 10.59 2.57 x 10<sup>-11</sup> M Acid 8.81 1.55 x 10<sup>-9</sup> M 5.19 6.46 x 10<sup>-6</sup> M Base 8.69 2.04 x 10<sup>-9</sup> M 5.31 4.88 x 10<sup>-6</sup> M Base 8.46 3.47 x 10<sup>-9</sup> M 5.54 2.88 x 10<sup>-6</sup> M Base

## pH and pOH - Ms. Mogck's Classroom

This also is an excellent representation of the concept of pH neutrality, where equal concentrations of [H<sup>+</sup>] and [OH<sup>-</sup>] result in having both pH and pOH as 7. pH+pOH=14.00. pH=14-pOH. pH=14-pH. Example 1: What is the pH of a solution that has a pOH of 12.40? Keep in mind that the relationship of pH and pOH equals 14. pH+pOH = 14.00. pOH = 12 ...

## How to Calculate pH in Chemistry | Albert.io

New York: Prentice Hall. ISBN 0-582-22628-7. pH, pKa, Ka, pKb, and Kb Explained. How to Find pOH in Chemistry. How to Calculate the pH of a Weak Acid. What Does pH Stand For? pH Definition and Equation in Chemistry. pH and pKa Relationship: The Henderson-Hasselbalch Equation. Is Negative pH Possible?

## Here's How to Calculate pH Values - ThoughtCo

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