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CHEMISTRY Matter and Change . Section  
10.1 Measuring Matter Section 10.2  
Mass and the Mole Section 10.3 Moles of  
Compounds Section 10.4 Empirical and  
Molecular Formulas Section 10.5

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Formulas of Hydrates Exit CHAPTER  
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Chapter 12: States of Matter CHEMISTRY  
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Chapter 8: Covalent Bonding  
Chapter 9: Chemical Reactions  
Chapter 10: The Mole  
Chapter 11: Stoichiometry  
Chapter 12: States of Matter  
Chapter 13: Gases  
Chapter 14: Mixtures and Solutions

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Elena is a chemist, a scientist who studies matter and the changes that it undergoes. Carl is very dependable. His teachers and his parents know that he is reliable and can be trusted. An air of jubilation surrounded the members of the science team as they received their medals for first place in the national competition.

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10 161 Section 10.1 Measuring Matter  
page 320-324 Practice Problems pages  
323-324 1. Zinc (Zn) is used to form a  
corrosion-inhibiting surface on  
galvanized steel. Determine the number  
of Zn atoms in 2.50 mol of Zn. 2.50 mol  
Zn

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Chemistry: Matter and Change Chapter  
17 Name 17 CHAPTER Section 174  
continued Date Class STUDY GUIDE FOR  
CONTENT MASTE In the space at the left,  
write true if the statement is true; if the  
statement is false, change the italicized  
word or phrase to make it true 8 To  
determine the instantaneous rate, you  
must know the specific

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In this chapter you will learn- Concept of  
matter. Matter in our surroundings (full  
chapter). Matter in our surroundings  
Notes. Notes and explanation. Short



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Chemistry: Matter and Change Chapter 8  
44 . Name Date CHAPTER FOR Class  
Section 8.2 continued 9 What is the  
relationship between lattice energy and  
the strength of the attractive force  
holding ions in place? a. The more  
positive the lattice energy is, the greater  
the force. b. The more negative the  
lattice energy is, the greater the force.

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### **CHAPTER 1 Matter and Change**

CCHEMISTRYHEMISTRY MMATTER AND CHANGEATTER AND CHANGE Consultant Douglas Fisher, PhD

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